

# Making Solutions

⌘ How many grams of NaCl would you need to prepare 200.0 mL of a 5 M solution?

$$g = M \times L \times \text{molar mass}$$

$$g = (5\text{mol/L}) (0.2\text{L}) (58.44\text{g/mol})$$

$$g = 58.44 \text{ g}$$

